

Original Article

ELECTROCHEMICAL AND CHEMOMETRIC DETERMINATION OF DORZOLAMIDE AND TIMOLOL IN EYE DROPS USING MODIFIED MULTIWALL CARBON NANOTUBE ELECTRODE

HASSAN A. M. HENDAWY¹, HANAN M. ELWY¹, AMANY M. FEKRY²

¹National Organization for Drug Control and Research (NODCAR), P. O. Box 29, Cairo, Egypt, ²Cairo University, Faculty of Science, Department of Chemistry
Email: hanan18s@yahoo.com

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ABSTRACT

Objective: This work is focused on the construction of simple and sensitive electrochemical sensor for quantitative determination of dorzolamide (DOR) and timolol maleate (TIM). This method is based on the incorporation of multiwall carbon nanotubes (MWCNT) into the carbon paste electrode which improve the characteristics of the electrode.

Methods: The electrochemical response of modified electrode was based on voltammetric oxidation, using cyclic voltammetry (CV) and impedance spectroscopy (EIS). The structural morphology of the surface modified electrode was characterized by scanning electron microscope (SEM). Quantitative analysis for each of the two compounds in a mixture has been examined by using of chemometric tools for resolving overlapping signals. The prediction performance of the chemometric method was analyzed by principal component regression (PCR) and partial least square (PLS).

Results: Fractional factorial design was constructed from set of synthetic mixtures of two drugs in concentration ranges of 0.05 to 1.6 µg/ml for DOR and 1.5-20 µg/ml for TIM. Under optimum experimental conditions, DOR and TIM gave rectilinear response over the concentration range of 0.072-1.88 µg/ml and 1.16-20.84 µg/ml, respectively. The limit of detection (LOD) was found to be 0.098 and 1.025 µg/ml, for DOR and TIM, respectively. It found that the % of relative prediction error (RPE) was acceptable and satisfactory.

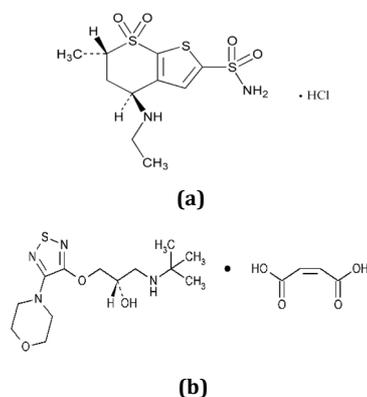
Conclusion: In these work, for the first time, a new voltammetric simultaneous method developed for a rapid and efficient determination of DOR and TIM from eye dropper sample at nano modified electrode with satisfactory results. These results indicate that MWCNT holds great promise in practical application.

Keywords: Dorzolamide, Timolol, Square-wave voltammetry, Partial least squares, Chemometric

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INTRODUCTION

Dorzolamide hydrochloride is a topical carbonic anhydrase inhibitor [1]. Inhibition of carbonic anhydrase-II decreases aqueous humor secretion. It is indicated for the reduction of elevated intraocular pressure in patients with open-angle glaucoma or ocular hypertension who are insufficiently responsive to beta blockers. Timolol maleate (TIM) is a non-selective beta adrenergic receptor blocker [2]. It is used to treat open-angle and occasionally secondary glaucoma through blockage of the beta receptors on the ciliary epithelium. Both DOR and TIM are used to lower pressure inside the eye by decreasing the production of natural fluids. The combined mixture of timolol maleate and dorzolamide hydrochloride is most advantageous than monotherapy. Their chemical structures are shown in Scheme 1.



Scheme 1: Chemical structures of (a) dorzolamide hydrochloride, [1], (b) timolol maleate [2]

The literature survey reveals that both dorzolamide hydrochloride and timolol maleate have been analyzed by various techniques. Several methods have been reported for the determination of DOR and TIM in biological fluids [3-6] by HPLC [7-10], spectrophotometry [11-13], capillary electrophoretic (CE) [14], and there are some methods reported for estimation of DOR in pharmaceutical formulations [8, 15-16]. Electrochemical detection is a very elegant method in the pharmaceutical analysis due to its high sensitivity, rapid response. Recently, CNTs are widely used in many types of sensors for determination of many compounds [17, 18]. CNTs are porous materials that offer high effective surface area, mass transfer, high electrical conductivity, and strong adsorption capacity resulting in increased sensitivity of sensors [19]. Thus, CNTs can powerfully enhance electron transfer reaction and improve sensitivity in electrochemistry. Consequently, CNTs have been explored as the most popular nano material used to prepare modified electrodes in the electro-analysis. Specially, the multiwall carbon nanotubes (MWCNTs) were thought to have much higher stability against electro-migration than other small metallic structures. MWCNT promote an effectual electronic transfer and are widely used in electro analytical chemistry for the sensitive detection of pharmaceutical compounds. Generally, there are two methods to prepare modified CNTs-based electrodes, in the first method, the CNTs is suspended on the surface of the solid electrodes, such as Pt and Au to make CNTs film modified electrodes [20-22]. The second method depends on mixing the CNTs with bonds such as mineral oil or nujol [23], to prepare modified paste electrodes. Differential pulse voltammetry (DPV) and square wave voltammetry (SWV) are widely used techniques in electrochemistry [24]. No electrochemical methodologies have been studied for simultaneous estimation of DOR and TIM, probably due to their similar electrochemical responses showing overlapping signals when using conventional electrodes. Due to the applicability of SWV

is ^{handicapped} when overlapping of peaks is high. Therefore, the analysis of voltammetric data for these signals is difficult and direct determination of analytes with overlapping voltammogram peaks is impossible. So recently, chemometric methods such as the multivariate curve resolution and partial least squares (MCR-PLS) [25-27] have been applied for the analysis of binary and multi components with overlapping results. The main advantages of these techniques are the higher speed processing data concerning the values of concentrations and voltammograms of compounds with strong overlapping. Our objective in the present work was to develop a new voltammetric method for a rapid and efficient determination of DOR and TIM in the ophthalmic pharmaceutical formulation using MWCNT modified carbon paste electrode.

MATERIALS AND METHODS

Chemicals and reagents

All reagents were of analytical grade and used without further purification. DOR and TIM standard supplied from National Organization for Drug Control and Research NODCAR, acetazolamide obtained from local market. Britton-Robinson (BR) buffer 0.04 M [28] was prepared by mixing an adequate amount of 0.2 M NaOH with different volumes of 0.04 M each of H₃PO₄, 0.04 M acetic acid and 0.04 M boric acid to obtain pH values varied from 2 to 10. Graphite powder (particle dimension 20 μm) and MWCNTs (purified to more than 95% C with a diameter of 6–9 nm and length of 5 μm) were obtained from Sigma (Egypt).

Standard solutions and working solutions

Stock solution of DOR and TIM (1.0 mg ml⁻¹) was prepared in bidistilled water and stored in a refrigerator at 4°C. Standard working solutions prepared daily by serial dilution of the stock standard solution.

Apparatus

All voltammetric measurements were performed using Metrohm electro analyzers model 797VA Computrace. These measurements were recorded by using VA Computrace version 1.3.1. The one compartment cell with the three electrodes consisted of the working electrode; a platinum wire was used as the auxiliary electrode. All the cell potentials were measured with respect to Ag/AgCl (3M KCl) reference electrode. Electrochemical impedance spectroscopic (EIS) measurements are performed using SP-150 potentiostat supplied with EC-Lab® software package. The impedance designs were recorded at the peak potential by applying 10 mV sinusoidal potential through a frequency range of 1.0 mHz to 100 kHz. A JENWAY 3510 pH meter with a glass combination served to carry out the pH measurements. All experiments were carried out at room temperature of 25±0.1°C, repeated for 3 times and reproducible results were obtained.

Preparation of carbon paste and MWCNT modified electrode

The unmodified Carbon paste electrode was prepared by mixing graphite powder 0.225 g with paraffin oil (seven to eight drops, ~0.2 ml) in a mortar. The carbon paste was packed into the hole of the electrode body and smoothed on a filter paper until its shiny appearance. The modified electrode was prepared by mixing graphite powder (0.465 g) and MWCNT (0.035g) were mixed well for 10 min with an appropriate amount (seven to eight drops, ~0.2 ml) of paraffin oil as pasting liquid in a mortar and wetted paste was obtained [29,30]. The carbon paste was packed into the hole of the insulin syringe body with diameter 3.0 mm. A new surface was obtained by removing excess paste out of the syringe tip and smoothed with a weighing paper until it has shiny appearance. Electrical contact was prepared by pushing a copper wire down into the syringe.

Analytical procedure

The prepared modified electrode with MWCNTs was activated in a 0.04M BR buffer as supporting electrolyte solution by CV sweeps between 0.0V and +1.3V potential range with a scan rate of 100 mVs⁻¹ until stable voltammograms were obtained (n = 10). When the current became steady, the modified electrode surface was washed

with distilled water. After each set of measurements, the electrode surface modified with MWCNTs was regenerated by successive cyclic voltammograms (n = 10) in the supporting electrolyte solution to remove charge impurities.

Chemometric methods

Chemometrics methods have played a vital role in facilitating the simultaneous analysis, PCR and PLS are factor analysis multivariate statistical tools which have many benefits and have been applied for analysis of multicomponent mixtures. In this paper only outline of the two chemometric methods will be provided.

Calibration and validation sets

A calibration set was outlined according to a fractional factorial design at several levels. PCR and PLS multivariate calibration methods involve constructing of the calibration set including a series of two drugs mixtures at five concentration levels. A calibration training set of twenty-five standard solutions was prepared in the concentration range of 0.05-1.6 and 1.5-20 μg/ml for DOR and TIM, respectively. The concentration set and its composition was presented in (table 4).

Model efficiency estimation

A model can be applied for analysis of mixture samples or not, model validation is potentially the common critical stride in the model building sequence. In order to estimate the performance of models, validation of each model was investigated for the prediction of validation and test sets, The relative root mean square error (%RMSE) was calculated for the models developed by using 3-5 latent variables for the calibration (%RMSEC), the prediction (%RMSEP) sets, and relative error of prediction (REP), for both DOR and TIM, respectively.

$$\text{RMSEP} = \frac{\sqrt{\sum_1^n (\text{Ypred} - \text{Yact})^2}}{n}$$

$$\text{REP}(\%) = \frac{100 * \sqrt{\frac{1}{n} \sum_{i=1}^n (\text{Ypred} - \text{Yact})^2}}{\text{Ymean}}$$

Where y_{act} and y_{pred} are actual and predicted concentrations of each component, respectively, and n is the number of samples in validation or test set.

RESULTS AND DISCUSSION

Carbon paste electrodes have several advantages of very low Ohmic resistance, low cost, very short response time, reproducibility of the preparation process, simplicity, cheap and quick preparation process which provide the possibility of measurements on small volumes. Therefore, the goal of this article is a fabrication of novel modified carbon paste electrodes for the determination of DOR and TIM in a pharmaceutical preparation.

Morphology of the different electrodes

Scanning electron microscope (SEM)

The SEM images of both CPE and MWCNT modified electrode are shown in fig. 1. The response of an electrochemical sensor was related to its physical morphology. Significant differences are observed in the surface structure of the two electrodes. The SEM image of the bare CPE is irregularly shaped graphite flakes and separated layers. While, the SEM image of MWCNT modified electrode shows flowery shaped SEM with a physical appearance and surface characteristics of, most of the particles homogeneously dispersed on the electrode's surface. This flowery like structure; displays a high surface area as shown in fig. 1.

Impedance of CPE and MWCNT modified

Impedance plots are shown as Nyquist plots (fig. 2 inset) for both the Bare and modified electrodes. It shows a semicircle corresponds to a charge transfer resistance at high frequency and a line corresponds to a diffusion process at low frequencies, respectively. The best model used to fit the experiments (fig. 2 inset) is a two time

constant model consists of R_s (solution resistance), R_1 , R_2 (resistances of inner and outer layers, respectively), [31] W (Warburg impedance) and C_1 , C_2 (capacitances of inner and outer layers, respectively). The data simulated to this equivalent circuit with a reasonable fit (2% average error). W is related to the linear region at low frequency (diffusion process) [32, 33], R_1 , R_2 are related to the semicircle at high frequency (charge transfer resistances of inner and outer layers, respectively). So, the mechanism of the reaction depends on both charge transfer and diffusion process. The fitting performed using EC-Lab® software [32, 34-36]. The impedance fitting parameters for the two electrodes are listed in (table 1). The MWCNT electrode shows a higher value for the inner capacitance ($1.7 \mu\text{F}$), Warburg impedance ($315.1 \Omega \text{ s}^{-1/2}$) and lower values for the charge transfer resistance (1500Ω) indicating a higher conductivity compared to the bare electrode (3000Ω). These results approve well the highest oxidation peak current obtained from CV's results to MWCNT electrode.

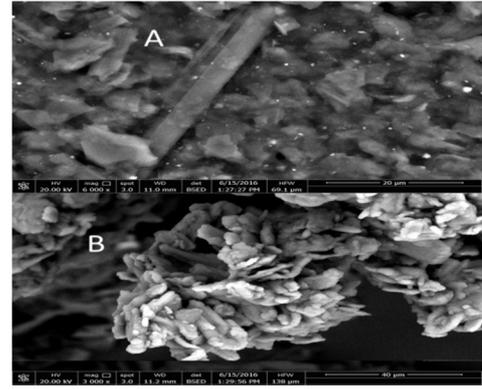


Fig. 1: SEM images of MWCNT (A) and CPE (B) electrodes

Table 1: Fitting parameters for bare and CNT electrodes

Electrode	$R_s (\Omega \text{ cm}^2)$	$C_1 (\mu\text{F cm}^{-2})$	$R_1 (\Omega \text{ cm}^2)$	$C_2 (\mu\text{F cm}^{-2})$	$R_2 (\Omega \text{ cm}^2)$	$W (\Omega \text{ cm}^2 \text{ s}^{-1/2})$
Bare CPE	400	0.4	3000	2.2	500	111
MWCNT	320.8	1.7	623	3.1	121.5	315.1

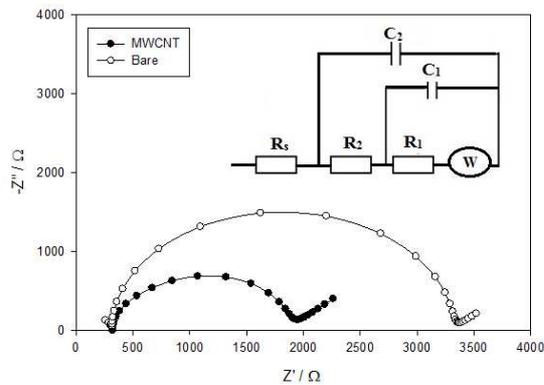


Fig. 2: Nyquist plots for both bare and MWCNT electrodes. Inset: Two time constant equivalent circuit model including Warburg diffusion impedance

Electrochemistry of dorzolamide HCl and timolol maleate

The voltammetric behavior of DOR and TIM in B-R buffer pH 6 at a scan rate 100 mVs^{-1} displayed a well-defined irreversible oxidation peak, and no reduction signal was observed in the reverse scan. The addition of MWCNT in the matrix of CPE was examined by comparing the voltammetric response of electrode. The sensitivity of the modified electrode was higher than bare electrode; it causes the increase of anodic peak current. The enhancement in current was attributed to the larger surface area of the modified electrode that improves the conductivity of the electrode. The suggested oxidation mechanism of DOR and TIM was presented in scheme 2.

Effect of pH

The influence of pH on the electrocatalytic oxidation of DOR and TIM was studied at the modified CPE using B-R buffer within pH range 3–10. It was observed that the oxidation peak potential shifted negatively with the increase in pH values that confirm the suggested mechanism in scheme 2. As depicted in fig. 3 the anodic peak potentials shifted gradually to less positive values with the increase in the solution pH, indicating the involvement of protons in the oxidation reactions. The highest oxidation current was observed at pH 6.0 for both DOR and TIM analysis. It was selected as the optimum pH for the determination of both drugs. The relation between peak potential and pH lead to a linear regression equations $E_p(\text{V}) = 1.4877 - 0.0603\text{pH}$ ($R^2=0.9827$) for DOR

$E_p(\text{V}) = 1.0885 - 0.0286\text{pH}$ ($R^2=0.9538$) for TIM.

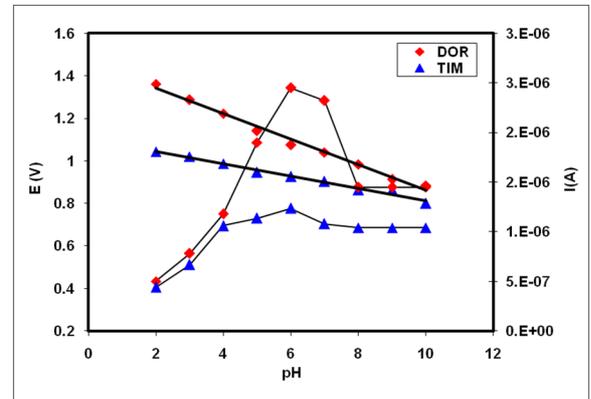
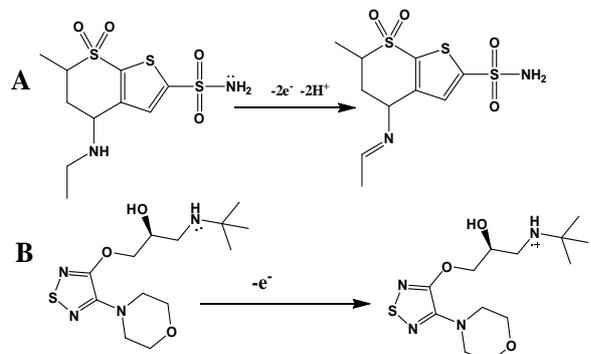


Fig. 3: Effect of pH on the peak potential for mixture solution of 20 mmol DOR and 10 mmol TIM in 0.04M BR at pH 6 at MWCNT modified electrode with scan rate of 100 mV s^{-1}

The result shows, that the slopes of DOR and TIM were -0.0603 and -0.0286 , respectively, [37] indicated that the system obeys Nernst equation. These results revealed that an equal number of participated protons and transferred electrons are involved in the oxidation case of DOR but not the same number in TIM. The proposed mechanism was a loss of lone pair of electrons for amine group with losing of two protons in the oxidation of DOR and loss of one electron in the case of TIM as depicted in scheme 2.



Scheme 2: Proposed oxidation mechanisms for the two drugs at MWCNT modified electrode

Effect of scan rate

Scan rate is one of the most important factors affecting the behaviour of electrochemical mechanism and kinetic of the cited drugs. The peak current grows with increasing scan rates, and plot of I_p versus $v^{1/2}$ gave linear relationships between the peak currents and the square root of the scan rate ($v^{1/2}$).

The oxidation peak potential shifts with increasing scan rates toward a more positive potential, indicating that oxidation process is irreversible. Scan rates were also studied to evaluate whether the process on MWCNT electrode was underneath diffusion or adsorption controlled. With increasing scan rate, the oxidation peak current of DOR and TIM increased, the relation of the square root of scan rate with the peak height of current exhibit a linear relationship in the range of 10-110 mV/s fig. 4 which represents a typical diffusion controlled process and equation can be expressed by

$I(A) = 27.322v^{1/2} + 0.2038$ $R^2 = 0.994$ (DOR), $I(A) = 39.341v^{1/2} + 0.2835$ $R^2 = 0.996$ (TIM)

Suggesting that the reaction is diffusion-controlled electrode reaction [38].

Linear calibration plot of single component and limit of detection

The SWV responses were studied at pH 6.0 on the MWCNT for a concentration range from 0.072–1.881 and 1.6 –20.84 $\mu\text{g/ml}$ of DOR and TIM, respectively, as shown in fig. 5(a and b). In addition to the

parameters of the calibration, the following conditions are used: voltage step 2 mV, amplitude 20 mV and frequency 50 Hz. Linear regression equation between I_p and concentration of drug can be expressed by the following equations: $I_p = 1.8053C + 0.1262$, ($R^2 = 0.9950$), $I_p = 0.0814C + 0.2966$, ($R^2 = 0.9976$), for DOR and TIM, respectively, the limits of detection (LOD) were found to be 0.017 and 0.932 for DOR and TIM, respectively. The relative parameters are summarized in (table 2).

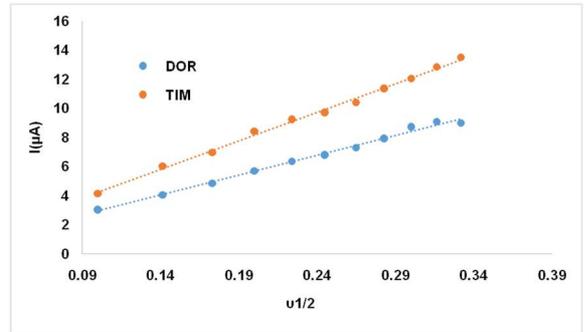


Fig. 4: linear relation between the anodic peak current of DOR, TIM and the square root of the scan rate (mVs^{-1}) at pH 6.0 for MWCNT with different scan rates (10-100 mVs^{-1})

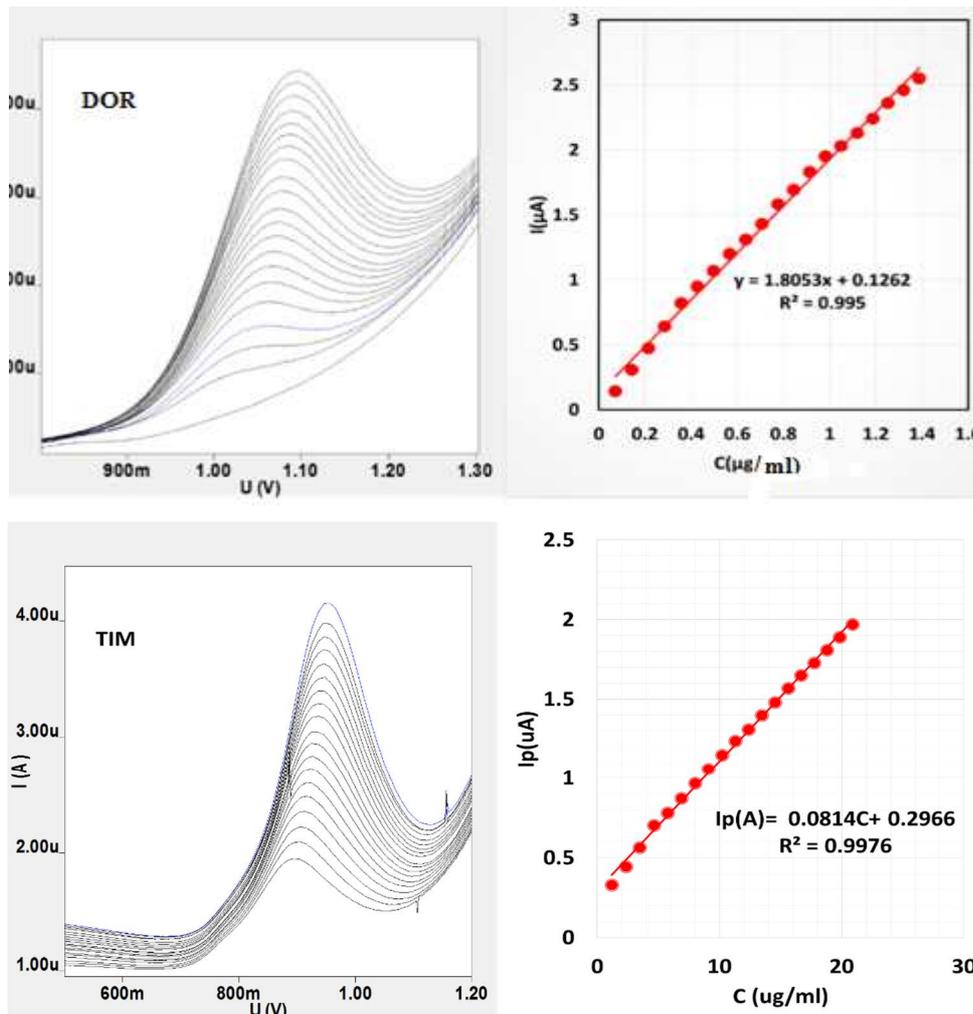


Fig. (5a, b): Square-wave voltammograms for calibration of DOR and TIM at pH 6.0 on the MWCNT from insert corresponds to the analytical curves obtained from voltammograms

Table 2: The parameters of linear calibration equation for each drug analyte

Parameter	TIM	DOR
Linear range ($\mu\text{g/ml}$)	1.16–20.84	0.072 – 1.88
Intercept (a)	0.2966	0.1262
Slope (b)	0.0814	1.8053
Correlation coefficient(R^2)	0.9976	0.995
Sa	0.0302	0.122
Sb	0.0253	0.01
Detection limit ($\mu\text{g/ml}$)	0.932	0.017

Sa: Standard deviation of the intercept, Sb: Standard deviation of the slope.

Effect of SWV

A study for Inferable from high affect ability and transcendent detachment properties from background current, SWV utilized to examine the dependencies relation between peak currents of DOR and TIM and their concentrations. The improvement of this system incorporates an advancement of instrumental parameters that influence current response of both drugs, such as pulse amplitude,

frequency and step increment. These parameters interrelate and influence the peak current. In a way that every parameter was changed while the others were kept constant. It was found that the peak currents of DOR and TIM increased simultaneously with the increasing of pulse amplitude from 10 to 140 mV joined by the widening peak widths at the same time. At the point when the pulse amplitude was higher than 60 mV, both oxidation peaks became much wider fig. 6.

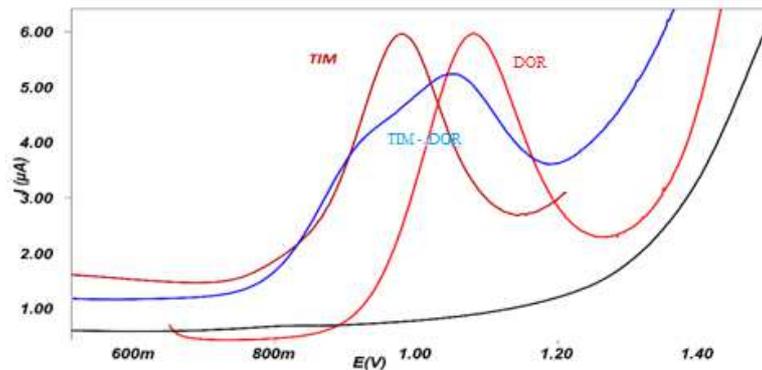


Fig. 6: Square wave voltammetric curves of 20 mmol DOR and 20 mmol TIM and for mixture solution 20 mmol DOR and 15 mmol TIM in 0.04M BR at pH = 6 buffer at MWCNT modified electrode

Quantitative analysis of individual drug was studied using SWV. The effect of the experimental parameters that affect the SWV response was evaluated for square wave voltammetric curves of 20 mmol DOR and 20 mmol TIM and for mixture solution 20 mmol DOR and 15 mmol TIM in 0.04 M BR at pH=6 buffer at MWCNT modified electrode (fig. 6) The obtained optimum values for these parameters are presented in (table 2).

Application of pharmaceutical formulations

The validity of the proposed voltammetric method was assessed for the determination of DOR and TIM in their pharmaceutical preparations. The mean recovery values were 99.95 and 100.02% with RSD 0.068 and 0.031% for DOR and TIM, respectively. The results of the proposed method are compared with those obtained from the reference methods [39] applying the Student's t-test and F-

test values did not exceed the theoretical values of 2.776 and 19 for t-and F-tests, respectively, (table 3). This comparison indicated that the proposed methods is not only as accurate as the reference methods, but it use simple reagent and apparatus; it is also applicable to a wide range of concentration besides being time-saving. The calculated t values (0.612 for DOR and 0.299 for TIM) were smaller than t Tabulated (2.776). The results are listed in (table 3). Hence, the proposed voltammetric method is adequate for the simultaneous determination of DOR and TIM in pharmaceutical formulations. The results suggested that MWCNT has high reproducibility and would be a useful sensor for quantitative analysis DOR and TIM. The robustness was examined by evaluating the variation of pH 6.5 instead of pH 6. The obtained results indicated that the proposed procedures for DOR and TIM assay can be considered robust.

Table 3: Application of the proposed and comparison methods for determination of DOR and TIM in its pharmaceutical preparations

Dosage form	DOR	TIM	Ref. method ^a or [39]	Ref. method TM [39]
Cosopt® dropper	99.8	100.03	99.98	99.97
Found %	100.1	99.97	99.95	99.99
	99.95	100.08	100.1	100.1
mean±SE	99.95±0.086	100.03±0.031	100.01±0.046	100.02±0.04
Standard Deviation	0.15	0.55		
Variance	0.025	0.303		
t-test	0.612	0.299		
F-test	3.57	1.617		
t-Tabulated	2.776			
F-Tabulated	19			

^aMean± standard deviation of three determinations, ^bThe tabulated values of t-and F-tests.

Comparisons of the data obtained for DOR and TIM determination using MWCPE with various methods, showed that our sensor is more sensitive than the reported HPLC methods: (120-480 and 360-1440 µg/ml) [40], (40-120 and 10-30 µg/ml) [41] and (2-20.5 and 4-45 µg/ml) [8] for TIM and DOR, respectively.

Prediction of synthetic mixtures of the drugs

In this work, the quantitative analysis of drug mixtures by SWV and multivariate calibration methods of PCR, and PLS was investigated. A set of 25 calibration samples was prepared using concentration values commensurate with the linear range thresholds (table 4).

Table 4: Concentration data of calibration set (C 1-C 14), and test set (T 15-T 25)

Sample	TIM (µg/ml)	DOR (µg/ml)	Sample	TIM (µg/ml)	DOR (µg/ml)
C 1	20.0	0.83	C14	20.0	1.60
C 2	10.8	0.05	T 15	15.4	1.21
C 3	6.10	1.21	T 16	20.0	0.05
C 4	6.10	0.44	T 17	1.50	1.21
C 5	20.0	1.21	T 18	20.0	0.44
C 6	6.10	0.05	T 19	15.4	0.83
C 7	15.4	0.05	T 20	10.8	1.21
C 8	10.8	0.83	T 21	1.50	1.60
C 9	15.4	1.60	T 22	10.8	1.60
C10	10.8	0.44	T 23	6.10	0.83
C11	1.50	0.05	T 24	15.4	0.44
C12	1.50	0.44	T 25	1.50	0.83
C13	6.10	1.60			

Method validation

Ten synthetic mixtures containing DOR and TIM in different concentrations levels were prepared as an independent validation set. The validation using PCR and PLS method has been done for obtaining reliable results of the analysis. The percentage recoveries and relative standard deviations were indicated in (table 5). The selection of the optimal number of latent variables for the PCR and PLS models is an essential step to avoid more noise added to the calibration model. A lot of methods are used for determining the optimal number of latent variables. In this study a cross-validation method was employed for the selection of the number of latent variables, using in the calibration set as in fig. 7 (a, b). A cross validation method leaving out one sample at a time was used. Anywhere selected samples are kept out of the calibration and used for prediction. This is repetitive till all samples have been kept out once. Validation residual variance can then be calculated from the prediction residuals. In segmented cross validation, the samples are divided into sub categories or "segments". One segment at a time is kept out of the calibration. There are as many calibration series as segments so that predictions can be investigated on all samples. The application of PCR and PLS models for prediction was quite successful. In mixtures containing both drugs, the relative mean error (RME) values were less than 10% and the recoveries were close to 100%; these recoveries were acceptable being in the range of 97.5–100.4%. The PCR and PLS models were applied to the drug mixtures in this study. The root means squares error of cross-validation (RMSECV) was computed as an investigative test for inspecting the errors in the predicted applications. Five and three LVs were found suitable for PCR and PLS, respectively, as in fig. 7 (a, b). The results predicted by the multivariate methods for the training set model are summarized in (table 5).

Interference studies

Under the experimental optimum conditions, potential interferences such as ions and organic compounds are investigated by addition of possible interference traces to mixture solution containing 20 mmol DOR and TIM. It was found that the common ions such as Na⁺, Pb²⁺, SO₄²⁻, CO₃²⁻ and NO₃⁻ had no interfere with DOR and TIM. No interfering peaks have been recorded around the peak potentials of DOR and TIM in 60-fold excess of common urinary compounds such as glucose, uric acid and ascorbic acid (signal change below 5%). As an illustrated example, we observe that uric acid oxidized at +0.5 to +0.7 V according to pH value. These facts prove the good selectivity of the proposed method and offer the promising possibilities for simultaneous determination of DOR and TIM.

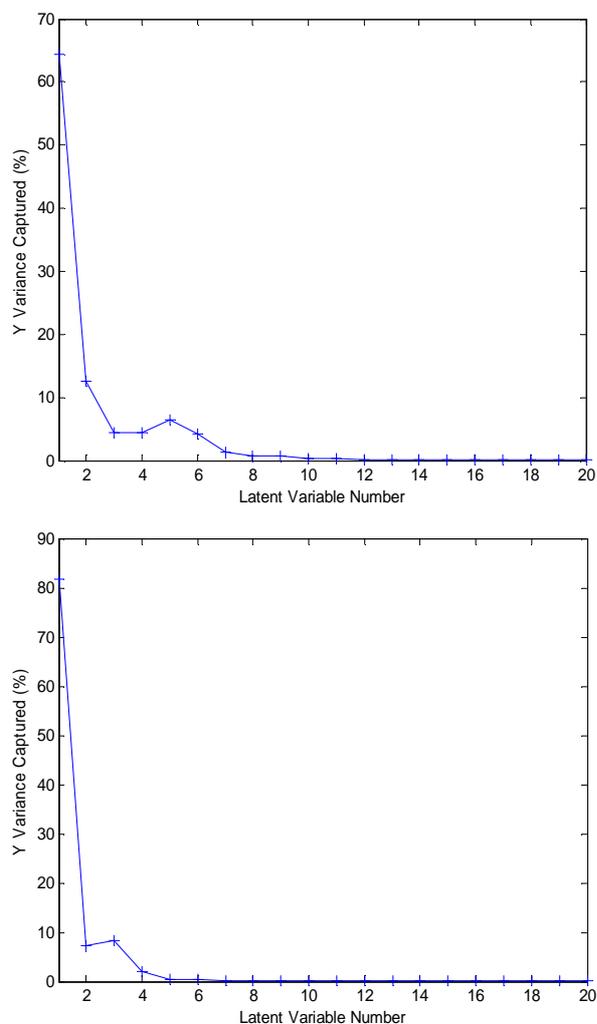


Fig. 7: RMSECV plot of the cross validation results of the calibration set as a function of the number of latent variables used to construct the PCR (a) and PLS (b) calibration

Table 5: Recovery results obtained for the synthetic mixtures for PCR and PLS methods

Mixtures added ($\mu\text{g/ml}$)		Recovery (%)		Recovery (%)	
		PCR		PLS	
DOR	TIM	DOR	TIM	DOR	TIM
15.4	1.21	96.96	96.85	99.47	99.14
20	0.05	98.99	105.71	101.8	102.01
1.5	1.21	85.73	118.57	100.71	99.09
20	0.44	102.98	103.08	99.45	98.51
15.4	0.83	102.99	104.07	97.50	98.44
10.8	1.21	89.67	88.96	99.49	102.62
1.5	1.6	109.9	108.2	102.48	100.09
10.8	1.6	99.73	93.59	101.28	102.34
6.1	0.83	97.18	91.43	99.50	102.30
15.4	0.44	89.51	95.37	102.54	101.92
1.5	0.83	98.87	95.97	98.77	98.07
Mean		97.5	100.16	100.27	100.41
RSD%		2.1	2.61	1.54	1.74
RMSECV		0.417	0.087	0.45	0.046
RMSEP		0.560	0.013	0.60	0.70

CONCLUSION

In the present paper, the electrochemistry combined with chemometrics tools has been developed an effective method for simultaneous determination of DOR and TIM. The two drugs exhibited a similar anodic behaviour; all are of irreversible character and pH-dependent. A strong and dramatic voltammetric overlapping observed for the simultaneous determination of these drug compounds when using suitable relabel electrodes. Chemometric tools PLS, PCR were successfully applied to resolve the overlapped voltammograms signals of these compounds in mixtures. It found that PLS and PCR improved the prediction results significantly, and yielded good REP values less than 10% for both the individual and overall estimates as well as good percentage recoveries. The proposed method applied successfully for the determination of the studied drugs in dosage forms without interference from excipients. The facts prove the good selectivity of the proposed method and offer the promising possibilities for simultaneous determination of DOR and TIM.

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CONFLICTS OF INTERESTS

All authors have none to declare

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